

Monazite, iron oxide and barite exsolutions in apatite aggregates from CCSD drillhole eclogites and their geological implications

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Abstract

We have identified abundant exsolutions in apatite aggregates from eclogitic drillhole samples of the Chinese Continental Scientific Drilling (CCSD) project. Electron microscope and laser Raman spectroscopy analyses show that the apatite is fluorapatite, whereas exsolutions that can be classified into four types: (A) platy to rhombic monazite exsolutions; (B) needle-like hematite exsolutions; (C) irregular magnetite and hematite intergrowths; and (D) needle-like strontian barite exsolutions. The widths and lengths of type A monazite exsolutions range from about 6–10 μm (mostly 6 μm) and about 50–75 μm , respectively. Type B exsolutions are parallel with the *C* axis of apatite, with widths ranging from 0.5 to 2 μm , with most around 1.5 μm , and lengths that vary dramatically from 6 to 50 μm . Type C exsolutions are also parallel with the *C* axis of apatite, with lengths of \sim 30–150 μm and widths of \sim 10 to 50 μm . Type D strontian barite exsolutions coexist mostly with type B hematite exsolutions, with widths of about 9 μm and lengths of about 60–70 μm . Exsolutions of types B, C and D have never been reported in apatites before. Most of the exsolutions are parallel with the *C* axis of apatite, implying that they were probably exsolved at roughly the same time. Dating by the chemical Th–U–total Pb isochron method (CHIME) yields an U–Pb isochron age of 202 ± 28 Ma for monazite exsolutions, suggesting that these exsolutions were formed during recrystallization and retrograde metamorphism of the exhumed ultrahigh pressure (UHP) rocks. Quartz veins hosting apatite aggregates were probably formed slightly earlier than 202 Ma. Abundant hematite exsolutions, as well as coexistence of magnetite/hematite and barite/hematite in the apatite, suggest that the oxygen fugacity of apatite aggregates is well above the sulfide-sulfur oxide buffer (SSO). Given that quartz veins host these apatite aggregates, they were probably deposited from SiO_2 -rich hydrous fluids formed during retrogression of the subducted slab. Such SiO_2 -rich hydrous fluids may act as an oxidizing agent, a feasible explanation for the high oxygen fugacity in convergent margin systems.

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1. INTRODUCTION

Mantle wedges above the subduction zone, and consequently arc magmas, are usually considerably more oxidized compared to mid-ocean ridge basalts (MORB) and ocean island basalts (OIB) (Ballhaus, 1993; Brandon and Draper, 1996; Parkinson and Arculus, 1999; Mungall, 2002; Sun et al., 2004b). Given that high oxygen fugacity is very important for Cu, Au ore genesis (Sillitoe, 1997; Mungall, 2002; Sun et al., 2004a; Liang et al., 2006), the mechanism that leads to oxidizing convergent margin systems is of wide interest. It is generally agreed that the high oxygen fugacity in the mantle wedge and arc magmas are due to the addition of subduction-released oxidizing fluids or melts (Brandon and Draper, 1996; Luhr and Aranda-Gomez, 1997; Parkinson and Arculus, 1999).

In general, C, H, S and Fe are elements, which are present in sufficient abundance and have variable oxidation states in the mantle, necessary to influence the redox states of the mantle (Mungall, 2002). Of the four elements, carbon and sulfur play very minor roles in oxidizing mantle rocks to oxygen fugacities above the SSO buffer (Brandon and Draper, 1996; Mungall, 2002). Water content is the most obvious difference between convergent margin environments and those for MORB and OIB: both mantle wedges and arc magmas have considerably more water than MORB and OIB due to the recycling of water-rich fluids released from subducting slabs (Wood et al., 1990; Arculus, 1994). The effects of water, however, depend on the physical separation of its decomposition products: H_2 and O_2 (Brandon and Draper, 1996), which is thought to be difficult in the mantle (Mungall, 2002). In contrast, Fe is believed to be the most important redox agent, and the only one that can introduce very high oxygen fugacity into the convergent margin system (Mungall, 2002). It was therefore argued that “if an arc magma has $\log fO_2 >$ SSO buffer, then it must contain a component of melted oceanic crust” (Mungall, 2002).

Here, however, we report abundant oxide exsolutions identified in apatite aggregates from eclogitic drillhole samples from the CCSD, which strongly suggests that solute-rich hydrous fluids can also act as an oxidizing agent that lead to high oxygen fugacity in convergent margin systems (Brandon and Draper, 1996).

2. GEOLOGICAL BACKGROUND AND SAMPLES

The Dabie–Sulu belt is the eastern part of the Qinling–Dabie orogenic belt in central China, which was formed by subduction (Sun et al., 2002c) and consequent collisions between different microplates (Li and Sun, 1996; Li et al., 1996; Zhang et al., 1996; Meng and Zhang, 2000; Sun et al., 2002b), followed by final continental collisions between the South and North China Blocks in the Triassic period (Li et al., 1993; Liou et al., 1996; Sun et al., 2002a,c; Zheng et al., 2003). It is, so far, the largest recognized ultrahigh pressure (UHP) metamorphic belt in the world. Studies show that the whole belt was once subducted down to more than 100 km depths (Xu et al., 1992; Jin et al., 1998; Yang and Jahn, 2000; Ye et al., 2000). Therefore, it has been a hot topic among geoscientists (Liou

and Zhang, 1996; Carswell et al., 2000; Li et al., 2000; Ye et al., 2002; Zheng et al., 2004; Xiao et al., 2006).

In order to investigate the evolution of the Qinling–Dabie orogenic belt, the CCSD project was carried out in the southern limb of the Su–Lu ultrahigh-pressure metamorphic terrain in Donghai County of Jiangsu Province (Xu et al., 1998; Xu, 2004; Zhang et al., 2005a) (Fig. 1). The drill core obtained from the main hole of the CCSD consists mainly of eclogites, ortho- and paragneisses, ultramafic rocks, and some schist and quartzite (Zhang et al., 2004, 2005c). The major rock-forming minerals are garnet, omphacite, phengite, quartz and kyanite, and the main accessory minerals are rutile, apatite and paragonite (Zhang et al., 2004). Apatite, which is one of the most common accessory minerals in the CCSD eclogites (Zhang et al., 2005b), is believed to have recorded subduction and exhumation history of the CCSD UHP rocks.

Apatite aggregates with oriented oxide inclusions have been found in many CCSD quartz vein samples (e.g., samples 2005011, 2005038, 2005049, 2003571, 2003588). The apatite samples reported here were collected from the CCSD main drill hole at the depth of 370 m. The sample number is 2003558 (B162R140P1g). The host rock is pyrite- and rutile-bearing quartz eclogite with quartz veins. Apatite occurs as massive aggregates to irregular veins with widths of 2–5 cm in quartz veins (Fig. 2a).

3. ANALYTICAL METHODS

Laser Raman analysis was done at the Guangzhou Institute of Geochemistry with a Renishaw RM 2000 Raman spectrometer with 5 mw of 514 nm Ar laser excitation at room temperature. The beam size for Raman spectroscopy was 2 μ m. Electron microprobe and scanning analyses were carried out in the Modern Analytical Center of Sun Yatsen University by using a JXA-8800R electron microprobe with an Oxford inca-300 energy dispersive X-ray spectrometer (EDS). The analytical accelerating voltage was 20 kV. The detection limit of the electron microprobe was 0.2% and resolution of the secondary electron images was 6 nm. Cobalt was used in calibration of peaks of EDS analysis.

CHIME dating of monazite exsolutions was performed by a JXA-8800M electron microprobe in the State Key Laboratory for Mineral Deposits Research at Nanjing University. The monazite grains for CHIME dating were selected under a polarized microscope and observed by a scanning electron microscope in order to avoid inclusions or fractures. UO_2 , ThO_2 and Pb-glass were used as standards for U, Th and Pb concentrations. The operating conditions were an accelerating voltage of 20 kV, electricity current of 1×10^{-7} A, and an electron beam diameter of 1 μ m. Measuring time and background were 100 and 50 s for U and Th, and 200 and 100 s for Pb. The detection limits for U, Th and Pb were 46, 60 and 46 ppm, respectively.

4. RESULTS

4.1. Mineralogy of oriented mineral inclusions

Electron microprobe analyses show that the apatite is fluorapatite in composition (Fig. 2b). Oriented mineral

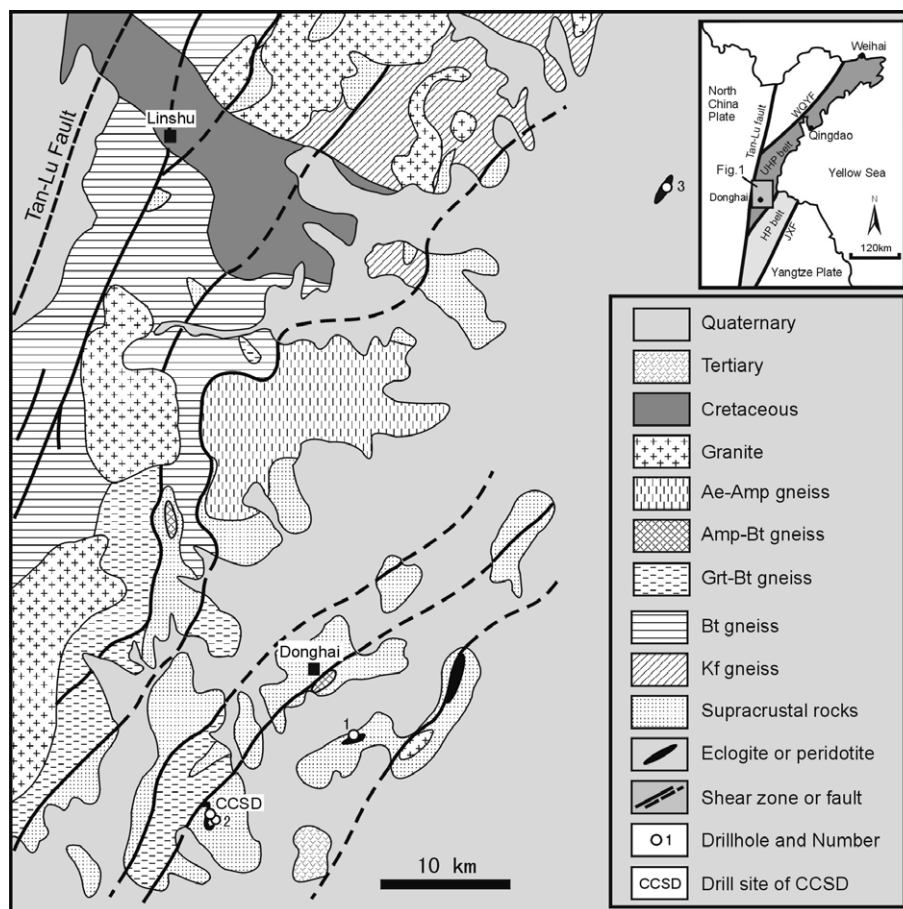


Fig. 1. Sketched geologic map of the Donghai area of the southern Sulu UHP belt, showing the location of the CCSD.

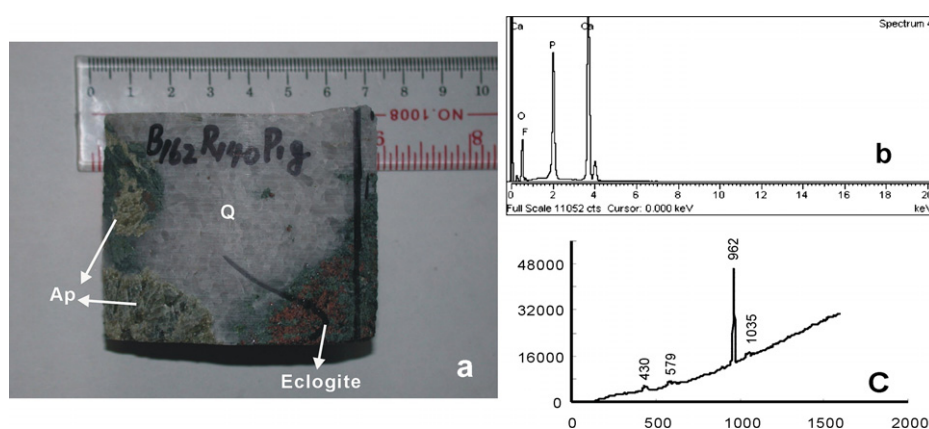


Fig. 2. Apatite aggregates (a) and their energy dispersive X-ray spectrometer (EDS) spectrum (b) and laser Raman spectrum (c) in CCSD eclogite. Ap-apatite; Q-quartz. The peaks at 430, 579, 962 and 1035 cm^{-1} represent apatite.

inclusions in the apatite can be classified into two types: transparent and opaque. The former are usually light yellow-color crystals of 6–10 $\mu\text{m} \times 50\text{--}70\text{ }\mu\text{m}$ in size, and unevenly distributed in apatite. The latter are mostly needle-like to irregular massive crystals. Both transparent and opaque inclusions are oriented needle-like grains, paralleling to the *C* axis of host apatite grains. Through

microprobe and laser Raman analyses, we have identified four types of oriented mineral inclusions in apatite:

[A.] Platy to rhombic monazite inclusions, which are mostly elongated parallel to the *C* axis of host apatite, have widths of 6–10 μm and lengths of 50–75 μm (Fig. 3a and b). Electron microprobe analyses

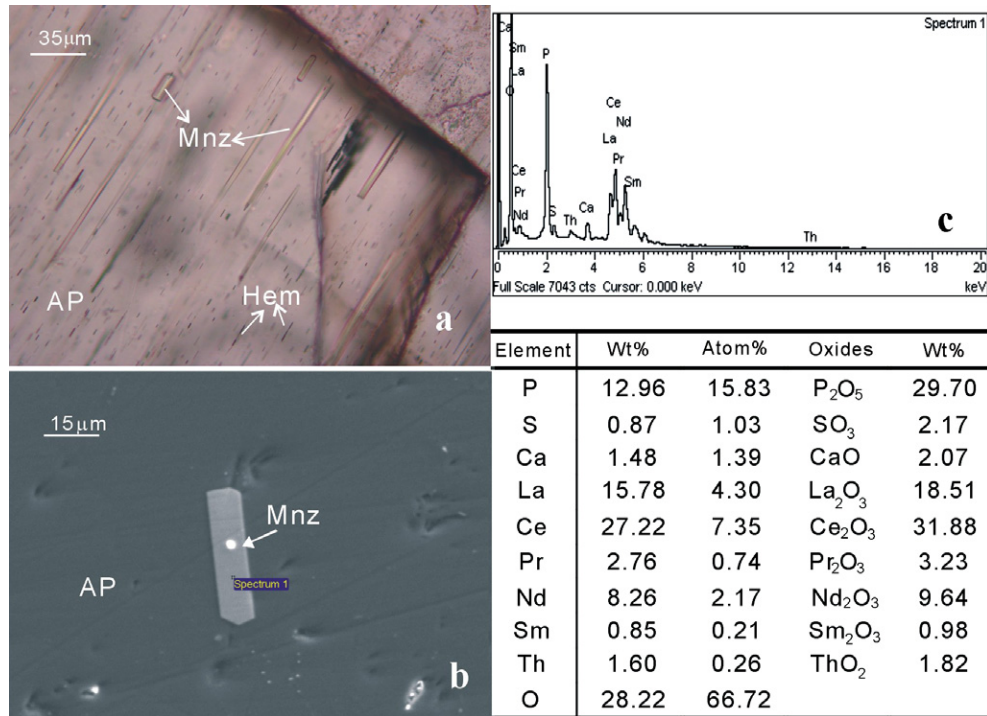


Fig. 3. Photomicrographs of oriented monazite inclusions (Type A) taken under microscope (a), BSE image (b), EDS spectrum (c) and chemical composition (3d). The mineral symbols: Mnz=monazite; Ap=apatite; Hem=hematite.

show that the monazites contain P, O, Th, S and high concentrations of light rare earth elements (LREE). The P and O contents as percentages of total atom weight are 15.83 and 66.72%, respectively, with P/O atomic ratios of approximately 1/4, indicating that they are LREE phosphate minerals. The total atomic percentage of LREE is 15.03% and $\text{LREE}/\text{PO}_4 \approx 1$. Therefore, the Type A inclusions are Th and S bearing monazites with a molecular formula of $(\text{LREE})\text{-PO}_4$. The ThO₂ concentration of this type of monazite is 1.82 wt% (Fig. 3), which is much higher than the usual 0.5% in monazite intergrown with apatite in typical metamorphic rocks (Andersson et al., 2002).

[B.] Needle-like hematite inclusions, which are also mostly elongated parallel to the *C* axis of host apatite, have widths of only 0.5–2 μm, with most around 1.5 μm, and lengths that vary dramatically from 6 to 50 μm (Fig. 4a). Electron microprobe analyses show that they are MgO-bearing (0.48%) iron oxides. Laser Raman analyses demonstrate that they are hematite with specific peaks of 226, 292, 408 and 653 cm⁻¹ (Fig. 4b).

[C.] Irregular magnetite and hematite intergrowths, which occur as massive opaque blebs with length of 30–150 μm and widths of 10–50 μm (Fig. 5a). Electron microprobe analyses show that they are iron oxides with 98.3 wt% Fe₂O₃, 0.66 wt% MgO and 1.04 wt%

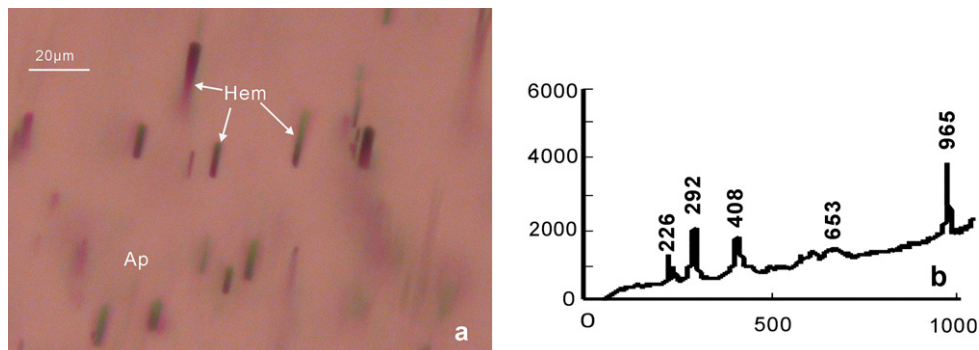


Fig. 4. Photomicrographs of oriented hematite inclusions (Type B) (a) and its laser Raman spectrum (b). The mineral symbols are same as Fig. 3. The peaks at 226, 292, 408 and 653 cm⁻¹ in 4b are specific peaks of hematite.

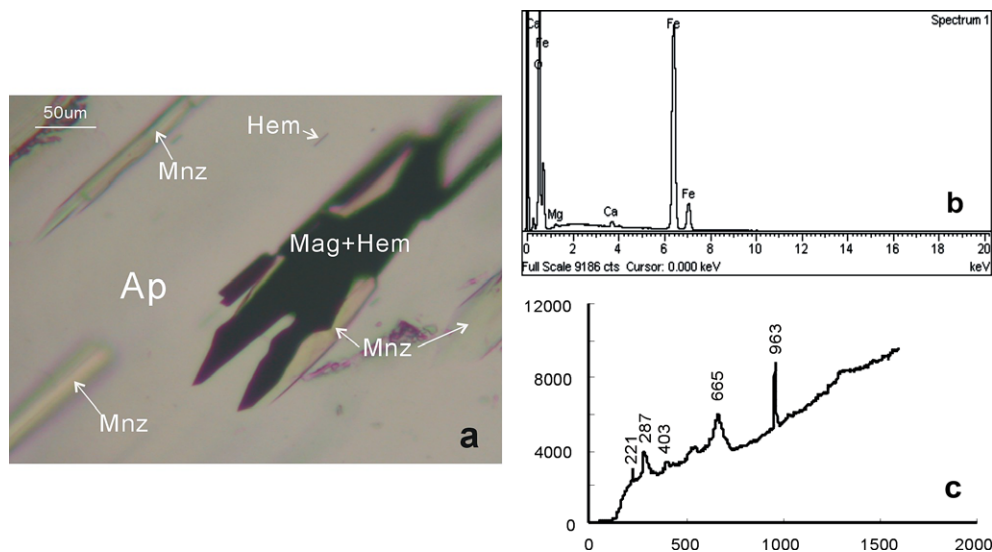


Fig. 5. Photomicrograph of magnetite and hematite intergrowth inclusions (Type C) (a), spectra of EDS (b) and laser Raman spectrum (c). The mineral symbols are same as Fig. 3. Mag-magnetite. The peaks in 5c at 221, 287 and 403 cm⁻¹ represent hematite, while the peak at 665 cm⁻¹ represents magnetite.

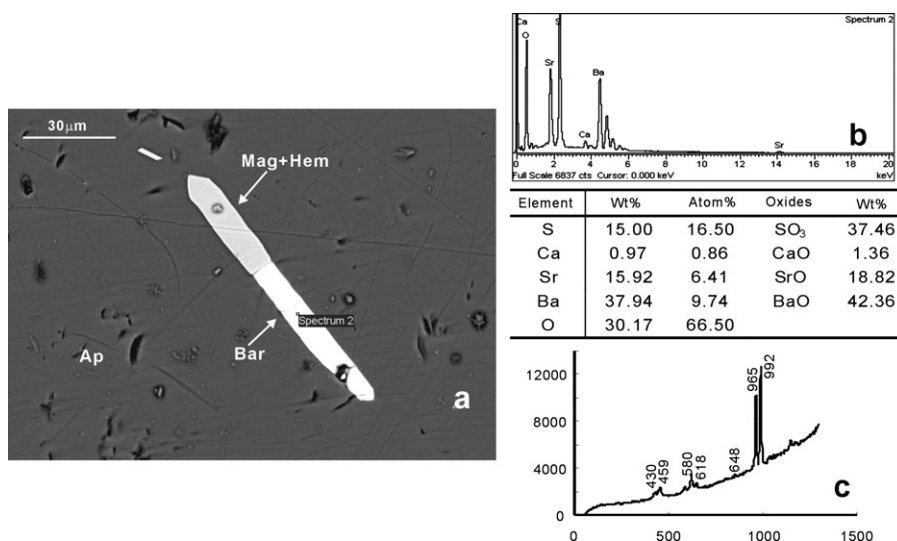


Fig. 6. BSE image of barite inclusion (Type D) (a), EDS spectrum and chemical composition (b) and laser Raman spectrum (c). The peaks at 992, 648, 618 and 459 cm⁻¹ in 6d represent barite. The upper part of the needle-like inclusions is a magnetite hematite intergrowth identified using laser Raman and electron microprobe analyses.

CaO. Laser Raman analyses demonstrate that they are hematite and magnetite intergrowths, the specific peaks for the former are 221, 287 and 403 cm⁻¹, while that for the latter is 665 cm⁻¹. The high CaO concentration is likely due to signals from the host.

[D.] Needle-like strontian barite inclusions have widths of about 9 μm and lengths of about 60 to 70 μm (Fig. 6a). Electron probe analyses show that they contain S, Sr, Ba, O and a small amount (0.86 wt%) of Ca. The atomic percentages of S and O are 16.5 and 66.5%, respectively, with S/O of about 1:4, confirming that they are sulfate minerals (Fig. 6b). In addition, the total atomic percentages of Ba and Sr

together are approximately the same as SO₄. Therefore, the type D mineral inclusions are strontian barite with a molecular formula of (Sr, Ba)SO₄. Laser Raman analyses demonstrate that they are barite with specific peaks at 992, 648, 618 and 459 cm⁻¹ (Fig. 6c).

4.2. CHIME dating of the monazite exsolutions

Monazite contains high amounts of radioactive elements, such as U and Th, with negligible common Pb (Williams et al., 1983; Corfu, 1988; Parrish, 1990; Williams and

Jercinovic, 2002), making it a suitable mineral for dating by CHIME (Suzuki and Adachi, 1991, 1998; Montel et al., 1996; Rhede et al., 1996; Cocherie and Albarede, 2001). Although the analytical error of CHIME ($\pm 10 \sim 30$ Ma) is larger than that associated with the sensitive high resolution microprobe (SHRIMP) method, the CHIME method is well-suited for dating monazite grains with a diameter of less than 20 μm , and particularly monazite exsolutions because the electron beam can be adjusted to 1–2 μm .

In this study, 51 monazite inclusions in apatite aggregates were analyzed using an electron microscope. The results are shown in Table 1 and plotted on Fig. 7. The isochron line gives an intercept of 0.0020 ± 0.0025 and a slope of 0.0085 ± 0.001 . The isochron age of the monazites is calculated with the CHIME software created by (Suzuki and Adachi, 1991, 1994), yielding a result of 202 ± 28 Ma, MSWD = 2.02.

5. DISCUSSION

For the following reasons, we propose that these oriented needle-like mineral inclusions in apatite formed through exsolution during decompression. First, all the inclusions are oriented roughly parallel to *C* axes of the host apatite. It is difficult to enclose such a large number of tiny minerals arranged in the same direction. Second, the inclusions are very small and needle-like; while their oxygen fugacities are far above those of host eclogites, none of them can survive without protection of apatite. Moreover, all the mineral inclusions are hosted only in apatite, with no inclusions found in any other minerals in the quartz veins. Given that apatite in quartz veins is younger than the host eclogite, it would be very difficult for the inclusions to be selectively captured by apatite.

Apatite has a specific hexagon structure with channels paralleling the *C* axis, and can thus hold a large number of elements with large radius ions under UHP; these elements can then be released during decompression and form exsolutions. Needle-like monazite inclusions in apatite have been found previously in Dabie–Sulu eclogites (Liou et al., 1998; Zhang and Liou, 1999). These authors proposed that the monazite inclusions may have exsolved from host apatite during decompression (Liou et al., 1998; Zhang and Liou, 1999). More recently, both monazite and sulfide (such as pyrrhotite and pyrite) inclusions in apatites from eclogites were attributed to exsolution during decompression (Hong et al., 2003; Zhu and Massonne, 2005). In this study, we found abundant mineral inclusions oriented along crystallographic axes in apatite aggregates associated with quartz veins in CCSD eclogites under microscope. Laser Raman and electron microprobe analyses demonstrated that the inclusions were monazite, iron oxides and barite. To our knowledge, this is the first report of iron oxide and barite exsolutions in apatite.

Abundant hematite exsolutions, as well as the coexistence of magnetite/hematite and barite/hematite exsolutions in apatite, suggest that the oxygen fugacity of apatite aggregates is well above the SSO buffer. By contrast, the host rock is pyrite- and rutile-bearing quartz eclogite, which has dramatically lower oxygen fugacity. These facts suggest that the

oxidizing components hosted in apatite originated in deeper levels and migrated to the current locations. Given that apatite aggregates are closely associated with quartz veins, these highly oxidizing components were probably closely associated with Si-rich hydrous fluids, which are very common fluids released from subducting slabs.

The origins of the quartz veins in the CCSD samples, and in Dabie–Sulu orogenic belt in general, remain uncertain. Early studies suggested that the quartz veins formed through an influx of later-stage external fluids (Li et al., 2001). Based on the lack of UHP minerals and isotopic data, it was later proposed that the quartz veins did not undergo UHP peak metamorphism, but formed mainly in P–T conditions between UHP and HP metamorphism of eclogites during exhumation of the subducted plate and before recrystallization of HP eclogites (Zheng, 2004), during retrograde metamorphism (Sun et al., 2006; Xu et al., 2006a,b). A recent study, however, reported the occurrence of $\text{N}_2\text{--CH}_4$ inclusions in quartz veins in CCSD eclogites, which indicates that at least some quartz veins in CCSD UHP rocks formed quite early and probably underwent UHP metamorphism (Xu et al., 2006a,b). Moreover, the lack of UHP minerals in quartz veins cannot be used to argue that these quartz veins formed at low pressure because, without protection (e.g., as inclusions in zircon or garnet), UHP minerals usually disappear during retrograde metamorphism. In fact, coesite pseudomorphs hosted in garnet and allanite in quartz veins have been found in many different locations, including CCSD samples (Zhai et al., 2005).

For Dabie UHPM rocks, three episodes of eclogite-facies metamorphisms, at 254–239, 238–230 and 218–206 Ma, respectively, have been identified based on SHRIMP dating and multiphase metamorphic textures and index mineral inclusions within zircon (Liu et al., 2006a,b). Based on evolutionary processes and geochronology, Zhang et al. (2005c) established the UHPM P–T–t path of the Dabie–Sulu belt and proposed that this area experienced epidote amphibolite facies prograde metamorphism (stage I), UHPM (stage II), coesite eclogite facies retrograde metamorphism (stage III), quartz eclogite retrograde metamorphism (stage IV), granulite facies retrograde metamorphism (stage V), amphibolite facies retrograde metamorphism (stage VI), epidote amphibolite facies retrograde metamorphism (stage VII), and greenschist facies retrograde metamorphism (stage VIII) (Zhang et al., 2005c). Based on previous results in geochronology and metamorphic petrology, the age of 202 ± 28 Ma for monazite exsolutions in the apatite aggregates implies that the formation of apatites occurred from quartz eclogite retrograde metamorphism (stage IV) to amphibolite facies retrograde metamorphism (stage VI) during exhumation of the subducted plates. The UHPM rocks were raised rapidly from a depth of 100–30 km at a rate of about 4 km/Ma, resulting in a dramatic decline in pressure, coupled with a slow decrease in temperature with a cooling rate of about 8 $^{\circ}\text{C}/\text{km}$, and simultaneous dehydration and decomposition of minerals (Xu et al., 2005; Zhang et al., 2005c). Some of the quartz veins may have formed as a consequence of this process. Overall, retrogression from eclogite to amphibolite metamorphism requires the addition of water. In other

Table 1

Electron microprobe analytical data of monazite exsolutions in apatite aggregates (wt%) and their apparent age (t, Ma)

Analytical No.	ThO ₂	UO ₂	PbO	Appearance age (t, Ma)	ThO ₂ *
1	3.103	0.157	0.027	177.42	3.609
2	2.915	0.171	0.034	232.16	3.469
3	1.901	0.107	0.021	221.4	2.247
4	3.121	0.2	0.028	176.34	3.766
5	1.992	0.082	0.02	210.01	2.257
6	2.645	0.189	0.035	254.3	3.258
7	1.097	0.066	0.014	252.79	1.311
8	2.683	0.137	0.029	219.78	3.126
9	1.129	0.082	0.013	220.89	1.394
10	2.962	0.178	0.032	214.34	3.538
11	2.055	0.14	0.026	245.38	2.509
12	2.073	0.103	0.023	226.41	2.406
13	1.179	0.092	0.015	240.48	1.477
14	1.737	0.152	0.019	202.12	2.228
15	1.379	0.063	0.013	194.76	1.582
16	1.431	0.098	0.021	283.93	1.75
17	2.583	0.139	0.025	195.48	3.032
18	1.388	0.074	0.012	174.98	1.627
19	1.98	0.073	0.017	181.97	2.215
20	2.078	0.113	0.029	280.56	2.445
21	2.036	0.142	0.034	321.54	2.499
22	1.938	0.132	0.018	180.59	2.364
23	4.798	0.363	0.048	190.63	5.97
24	2.798	0.111	0.031	232.54	3.157
25	1.899	0.094	0.019	204.44	2.203
26	1.641	0.269	0.024	226.36	2.512
27	2.342	0.103	0.024	212.6	2.675
28	2.269	0.091	0.024	221.81	2.563
29	2.537	0.127	0.032	256.85	2.949
30	3.689	0.214	0.047	253.84	4.383
31	2.293	0.102	0.022	198.86	2.622
32	2.581	0.108	0.034	274.39	2.932
33	2.553	0.099	0.029	238.98	2.874
34	1.912	0.106	0.022	231.06	2.255
35	4.285	0.252	0.05	232.16	5.101
36	1.979	0.127	0.022	218.12	2.39
37	1.259	0.06	0.017	276.63	1.454
38	2.14	0.107	0.026	247.53	2.487
39	2.547	0.13	0.032	255.16	2.969
40	1.962	0.09	0.018	189.47	2.252
41	1.431	0.105	0.021	280.32	1.772
42	1.007	0.072	0.013	248.13	1.24
43	1.897	0.122	0.022	227.38	2.292
44	1.191	0.08	0.013	212.53	1.45
45	1.642	0.099	0.023	277.11	1.964
46	1.282	0.103	0.019	278.04	1.617
47	1.149	0.085	0.016	265.73	1.425
48	1.528	0.092	0.02	259.18	1.826
49	1.54	0.1	0.018	228.77	1.864
50	1.73	0.132	0.024	263.16	2.158
51	1.483	0.257	0.021	215.03	2.314

Note. $\text{ThO}_2^* = \text{ThO}_2 + \frac{\text{UO}_2 \times W_{\text{Th}}}{W_{\text{U}}(\exp(\lambda_{232}t) - 1)} \times \frac{\exp(\lambda_{232}t) + 137.88 \exp(\lambda_{238}t)}{1} 38.88 - 1$, where W is molecular weight of each oxides, $W_{\text{Th}} = 264$, $W_{\text{U}} = 270$. Monazite contains high Th and its radiogenic Pb is mainly ^{208}Pb , thus $W_{\text{Pb}} = 224$. Radioactive decay constants (λ) of ^{232}Th , ^{238}U and ^{235}U are $4.9475 \times 10^{-11} \text{ a}^{-1}$, $1.55125 \times 10^{-10} \text{ a}^{-1}$ and $9.848 \times 10^{-10} \text{ a}^{-1}$, respectively.

words, the whole retrograde metamorphism process usually forms more hydrous minerals than those being decomposed. In contrast, hydrous fluids released from subducting slabs are usually Si-rich. Therefore, no matter what pressure the quartz veins formed at, the source materials were supplied through subduction.

We propose that apatite aggregates were probably formed at very high pressures, so that they contained high concentrations of oxides, and exsolutions occurred when pressure diminished during retrograde metamorphism. Impurity elements incorporated in apatite at high pressure became unstable and were expelled from the

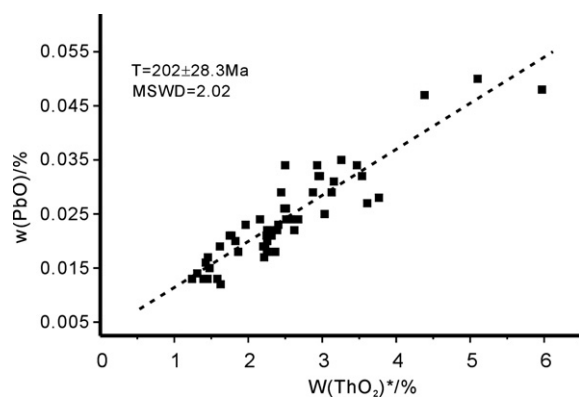


Fig. 7. Plot of ThO_2^* vs. PbO for monazite inclusions in apatite.

host minerals. Combination of the expelled LREE with P produced monazite exsolutions, while Sr and S combined to form strontian barite, and iron oxide formed magnetite and hematite.

Considering that apatite aggregates with oriented oxide inclusions have been found in many CCSD samples (e.g., samples 2005011, 2005038, 2005049, 2003571, 2003588), the appearance of apatite aggregates with a large amount of iron oxides and sulfates in the CCSD quartz veins suggests the formation of highly oxidizing SiO_2 -rich hydrous fluids during the subduction process. This kind of fluid is very likely to have contributed to the generation of high oxygen fugacity in convergent margin systems, mantle wedges, and arc magmas. These highly oxidizing components suggest that a slab-derived melt is not always necessary to produce arc magmas with high oxygen fugacity (above the SSO buffer). Highly oxidizing hydrous fluids may have played a more important role than previously recognized.

Note that the host eclogites are much more reducing compared to fluids, so a considerable amount of oxidizing components in fluids might have been lost through interactions with the wall rocks during ascent, resulting in the preservation of only the oxidizing components enclosed in apatite aggregates.

6. CONCLUSIONS

Apatite aggregates in quartz veins from CCSD drill hole samples were most likely formed through deposition of slab-derived Si-rich fluids. Abundant hematite exsolutions, as well as the coexistence of magnetite/hematite and barite/hematite assemblages in apatite, show that the oxygen fugacity of apatite aggregates is well above the SSO buffer. Consequently, slab-derived fluids can be highly oxidizing. This kind of fluid is very likely one of the efficient agents that caused the high oxygen fugacity in convergent margin systems.

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